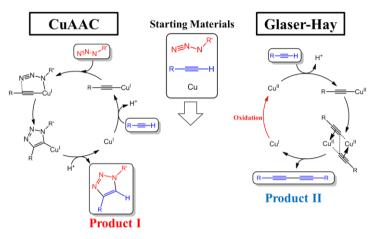
## Redox-Controlled Dual-Reactivity of Alkyne Catalyzed by Copper Complex

○<u>Yuya Nitta</u>, Kazuyuki Kubo, Shoko Kume and Tsutomu Mizuta Graduate School of Science, Hiroshima University

## ~Introduction~

Scheme 1. Glaser-Hay Coupling and CuAAC

system The converting an electric signal into a molecular framework can be а basic nanoscale tool to construct a selected structure at a desired position in molecular scale. We aimed to construct such a system from a redox-active catalyst which undergoes different coupling depending reactions on its oxidation state. There are two reactions of terminal alkyne with

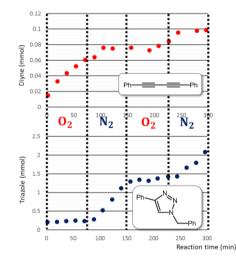


copper catalysts. Glaser-Hay Coupling is an oxidative reaction which affords 1,3-diynes by the homocoupling of terminal alkynes, and a Cu(II)/Cu(I) re-oxidation step by oxygen complete the catalytic cycle. Azide-Alkyne Cycloaddition is the reaction which affords 1,2,3-triazole from terminal alkynes and azides, without any oxidant. Thus, Glaser-Hay Coupling or CuAAC process can be selected from the same solution if we switch the oxidation process (**Scheme 1**).

## ~Experimental and Results~

The solution containing phenylacetylene, benzylazide (0.73 equiv.), copper(tmeda) complex (0.8 mol%) and TMEDA (= N, N, N', N'-tetramethylethylenediamine) under oxygen atmosphere formed diyne, however, a subproduct also formed through the oxidation of an intermediate from the CuAAC catalytic cycle. Decreasing the additional TMEDA from 18 mol% to 0.8 mol% resulted in suppression formation of the subproduct. of the Furthermore, the formation of divne or triazole can be switched several times in presence or absence of oxygen, respectively (Figure 1).

In the formation process of 1,3-diyne, the



**Figure 1.** Product formation profile from a solution containing 0.8 mol% CuCl2(tmeda), 0.8 mol% TMEDA, DME:MeOH(1:1), phenylacetylene, 0.73 eq. benzylazide, *o*-terphenyl as standard, under  $O_2$  or  $N_2$  atmosphere.

oxidation process by oxygen could be replaced with anodic oxidation. The reaction proceeded with almost quantitative current efficiency, but it decelerated in the course of the reaction, due to a disproportionation of the solution within the separated electrolysis cell. The application of the anodic oxidation to switch the dual-reactivity is now under investigation.